



## Cylinder Gas China Factory High Purity manufacture semiconductors gas Boron trichloride

### Our Product Introduction

#### Basic Information

- Place of Origin: China
- Brand Name: CMC
- Certification: COA
- Model Number: Bcl3
- Minimum Order Quantity: 1kg
- Price: US \$200-2000/pc
- Packaging Details: Cylinder
- Delivery Time: 15 days
- Payment Terms: L/C, T/T
- Supply Ability: 30,000tons/year



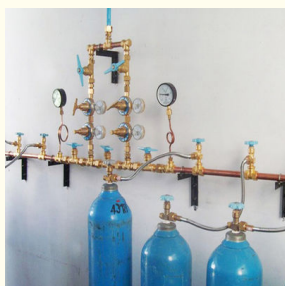
#### Product Specification

- Product Name: Boron Trichloride
- Cylinder Pressure: 15MPa/20MPa
- Cylinder Standard: GB/ISO/DOT
- Valve: Cga660
- Appearance: Colorless
- Transport Package: 40L/47L/50L
- Specification: 40L/47L/50L
- Trademark: CMC
- Origin: China
- HS Code: 2812191090
- Supply Ability: 300, 000tons/Year
- CAS No.: 10294-34-5
- Formula: Bcl3
- EINECS: 233-658-4
- Constituent: Industrial Pure Air



Boron Trichloride

#### More Images



## Product Description

### Product Description

Boron trichloride (BCl<sub>3</sub>) is a chemical compound composed of one boron atom and three chlorine atoms. It is a colorless gas with a pungent odor.

BCl<sub>3</sub> is known for its Lewis acidity, meaning it readily accepts electron pairs from other molecules or ions. Here are some key points about BCl<sub>3</sub>:

Chemical Formula: BCl<sub>3</sub>

Molecular Weight: 117.17 g/mol

Physical Properties: Boron trichloride is a gas at room temperature and atmospheric pressure. It has a boiling point of -107.8 degrees Celsius (-162.0 degrees Fahrenheit) and a melting point of -107.8 degrees Celsius (-162.0 degrees Fahrenheit).

Odor: BCl<sub>3</sub> has a sharp and irritating odor.

Lewis Acidity: BCl<sub>3</sub> is a Lewis acid, meaning it can accept an electron pair from a Lewis base to form a coordinate covalent bond.

Reactivity: Boron trichloride reacts with water to form boric acid (H<sub>3</sub>BO<sub>3</sub>) and hydrochloric acid (HCl).

$\text{BCl}_3 + 3\text{H}_2\text{O} \rightarrow \text{H}_3\text{BO}_3 + 3\text{HCl}$

It also reacts with alcohols to form alkyl chlorides:

$\text{BCl}_3 + 3\text{ROH} \rightarrow \text{B(OR)}_3 + 3\text{HCl}$

Uses: BCl<sub>3</sub> is primarily used as a reagent in organic synthesis, particularly in the production of boron compounds. It is also utilized in the manufacture of semiconductors and as a catalyst in various chemical reactions.

Safety Precautions: Boron trichloride is a toxic and corrosive substance. It can cause severe burns and eye damage upon contact. Proper safety measures, including the use of protective equipment, should be followed when working with BCl<sub>3</sub>.

#### Basic Info

Transport Package:	40L/47L/50L	Melting Point	-107.3°C
Trademark:	CMC	Boiling Point	12.5°C
Specification	99.90%	Production Capacity	300,000tons/Year
Cylinder Pressure	12.5MPa/15MPa/20MPa	Valve	Cga660
Appearance	Colorless Fuming Liquid or Gas with a Pungent Density		1.35 Kg/M3

#### Product Description

##### Specification:

Dot Class: 2.3  
State: Liquid  
Purity: 99.9%  
UN NO: UN1741  
CAS NO: 10294-34-5  
Grade Standard: Industrial Grade

Specification	99.9%
Chlorine	≤ 10 ppm
Silicon Tetrachloride	≤ 300 ppm

Cylinder Specifications	Contents
Cylinder Capacity	Valve Weight
47L	CGA 660 50 kgs

#### Detailed Photo



Company Profile



Shanghai Kemike Chemical Co., Ltd is staffed by trained personnel, combine many years experience in Gas industry .We supply cylinder gas, electronic gas, etc ., and the gas holder, panel, valves and fittings and other equipment, parts and engineering services to our customers in China and worldwide; The products are involved in various industrial fields, such as semiconductor chip, solar cell, LED, TFT-LCD, optical fiber, glass, laser, medicine , etc., Our mission is to partner with our global customers to provide support, solutions and quality products that are innovative, reliable, and safe. Our products mainly include: H<sub>2</sub>, O<sub>2</sub>, N<sub>2</sub>, Ar, CO<sub>2</sub>, propane, acetylene, helium, laser mixed gas, SiH<sub>4</sub>, SiH<sub>2</sub>Cl<sub>2</sub>, SiHCl<sub>3</sub>, SiCl<sub>4</sub>, NH<sub>3</sub>, CF<sub>4</sub>, NF<sub>3</sub>, SF<sub>6</sub>, HCl, N<sub>2</sub>O, doping mixed gas (TMB, PH<sub>3</sub>, B<sub>2</sub>H<sub>6</sub>) and other electronic gases.

SiCl <sub>4</sub>	NH <sub>3</sub>	NH <sub>3</sub>	CH <sub>3</sub> F	SiH <sub>4</sub>	Kr	H <sub>2</sub> S	WF <sub>6</sub>	F <sub>6</sub> +Cl <sub>2</sub>
4MS	C <sub>3</sub> F <sub>8</sub>	C <sub>3</sub> F <sub>8</sub>	TEOS	CH <sub>4</sub>	PH <sub>3</sub>	SF <sub>6</sub>	C <sub>2</sub>	HCl+Ne
CF <sub>4</sub>	C <sub>4</sub> F <sub>8</sub>	SiH <sub>2</sub>						TMB+H <sub>2</sub>
SiF <sub>4</sub>	C <sub>3</sub> H <sub>8</sub>	Cl <sub>2</sub>						He +As
BBr <sub>3</sub>	C <sub>3</sub> H <sub>6</sub>	DCE						Ge+Se
POCl <sub>3</sub>	N <sub>2</sub>	SO <sub>2</sub>						D+B
BCl <sub>3</sub>	D <sub>2</sub>	CO <sub>2</sub>						CO+NO
SiHCl <sub>3</sub>	CH <sub>2</sub> F <sub>2</sub>	HF						Ar+O <sub>2</sub>
TMAI	DMZn	DEZn						Xe+NO
AsH <sub>3</sub>	C <sub>2</sub> H <sub>4</sub>	C <sub>2</sub> H <sub>2</sub>	HBr	COS	Ar+O <sub>2</sub>			
GeH <sub>4</sub>	C <sub>2</sub> H <sub>6</sub>	B <sub>2</sub> H <sub>6</sub>	H <sub>2</sub> Se	GeCl <sub>4</sub>	Xe+NO			



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